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1. What is the number of atoms of oxygen in 2.0 mol of hydrated sodium carbonate, $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$? Avogadro's constant, L or N_A $23: 6.02 \times 10^{23} \text{ mol}^{-1}$ A. 6 B. 26 C. 24×103.6 D. 24×101.6
2. What is the volume, in cm^3 , of the final solution if 100 cm^3 of a solution containing 1.42 g of sodium sulfate, Na_2SO_4 , is diluted to the ...

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1.25 cm^3 1.0 mol dm^{-3} 25 cm^3 1.0 mol dm^{-3} 50 cm^3 1.0 mol dm^{-3} 50 cm^3 1.0 mol dm^{-3} y
What is the value of y ? A. 1 B. 2 C. 4 D. 15. What is the value of ΔH for the exothermic

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reaction represented by the diagram below? Enthalpy x y z Reactants Products Reaction coordinate
A. $y - z$ B. $z - y$ C. $x - z$ D. $z - x$

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1 1 H 1.01 3 Li 6.94 11 Na 22.99 19 K 39.10 37 Rb 85.47 55 Cs 132.91 87 Fr (223) Atomic number
Element Relative atomic mass 25 Mn 54.94 43 Tc 98.91 75 Re 186.21 61 Pm 146.92 93 Np (237)

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Some problems may include the stoich question (the volume one) on Topic 1, the shape of that weird molecule on Topic 4, and energetic manipulations on Topic 5. But overall, this paper was okay though. Past papers really helped. P.S. Topic 2 question on the number of electrons in NUCLEUS was LIT. Can you imagine how the IB tricked us so harshly lol.

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